

Stoichiometry Practice Test Answers

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Answers: Moles and Stoichiometry Practice Problems 1) How many moles of sodium atoms correspond to 1.56×10^{21} atoms of sodium? $1.56 \times 10^{21} \text{ atoms Na} \times 1 \text{ mol Na} = 2.59 \times 10^3 \text{ mol Na}$ $236.022 \times 10 \text{ atoms Na}$ 2) Determine the mass in grams of each of the following: a. 1.35 mol of Fe $1.35 \text{ mol Fe} \times 55.845 \text{ g Fe} = 75.4 \text{ g Fe}$ 1 mol Fe b. 24.5 mol O

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Answers to Stoichiometry Practice Free Response Test: 1. $30.40 \text{ g N} \times 1 \text{ mol N} / 14.01 \text{ g N} = 2.17 \text{ moles N}$. $69.60 \text{ g O} \times 1 \text{ mol O} / 16.00 \text{ g O} = 4.35 \text{ moles O}$. $4.35 / 2.17 = 2.00$ $2.17 / 2.17 = 1.00$ So O is 2X more present than N. Empirical formula must be NO. 2. Mass.

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Chapter 12 Stoichiometry Test Answer Key

Practice Test Ch 3 Stoichiometry Name Per. Remember it is a MC test, use the answers ... Practice Test Ch3 Stoichiometry (page 3 of 3) 1. d It might be easiest to balance the equation with mostly whole numbers: $2 \text{ NH}_3 + \text{O}_2 \rightarrow 2 \text{ NO}_2 + 3 \text{ H}_2$... 7. c First you must realize this is a limiting reactant problem. Found: 3 Jan 2020 | Rating: 89/100

Stoichiometry Practice Problems Level 1 Answers

2O Then do some stoichiometry using "easy math" 16 g of methane (MM = 16) is 1 mole and 1 mole of methane will produce 1 mole of CO₂ = 44 g, and 2 moles of H₂O which is 36 g for a total of 80 g 4. d Balance: $\text{C}_3\text{H}_8 + 5 \text{ O}_2 \rightarrow 3 \text{ CO}_2 + 4 \text{ H}_2\text{O}$ 5. d Balance: $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$

Practice Test Ch 3 Stoichiometry Name Per

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Stoichiometry Practice Test - Answer Key. Back to the other Stoichiometry Practice Tests and other General Chemistry Practice Tests. The formation of NH₃ from N₂ and H₂ occurs in 85.0% yield. How many grams of ammonia would be experimentally obtained when 12 g of H₂ reacts with 20g of N₂?

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Chemical Stoichiometry Test Answers

Stoichiometry Practice Test Short Answer: Aluminum bromide can be prepared by the reaction of aluminum metal with bromine gas shown by the equation: $2 \text{Al} + 3 \text{Br}_2 \rightarrow 2 \text{AlBr}_3$
Now suppose that 5.6 mol of aluminum reacts with 4.4 mol of bromine. 1. Calculate the mass of aluminum bromide that can be produced from 5.6 mol of Al. 2.

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