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not do “ long division ” to try to get exact values. Remember it is a MC test, use the answers • Mark which questions you would like to “ go over ” when we get to school in September. 1. Balance the following equation:  $\_\_ \text{NH}_3 + \_\_ \text{O}_2 \rightarrow \_\_ \text{NO}_2 + \_\_ \text{H}_2\text{O}$  The balanced equation shows that 1.00 mole of  $\text{NH}_3$  requires  $\_\_ \text{mole(s)}$  of  $\text{O}_2$  a. 0.57 b. 1.25 c. 1.33

### Practice Test Ch 3 Stoichiometry Name Per

fewer steps are required to solve stoichiometry problems when the reactant is given in moles and the product is sought in moles which of the following mathematical expressions correctly states the relationship among percentage yield, actual yield, and theoretical yield %yield= (act. yield/theo. yield)  $\times$  100

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## Chapter 9 Stoichiometry Test Answers

Practice Problems (Chapter 5): Stoichiometry CHEM 30A Part I: Using the conversion factors in your tool box g A mol A mol A 1. How many moles CH<sub>3</sub>OH are in 14.8 g CH<sub>3</sub>OH? 2. What is the mass in grams of 1.5 x 10<sup>16</sup> atoms S? 3. How many molecules of CO<sub>2</sub> are in 12.0 g CO<sub>2</sub>? 2 4. What is the mass in grams of 1 atom of Au? KEY Tool Box: To ...

## Practice Problems (Chapter 5): Stoichiometry

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## Chemistry Matter And Change Chapter 12 Stoichiometry Study ...

Stoichiometry Chapter 3! Stoichiometry: Calculations with Chemical Formulas and Equations. Stoichiometry Anatomy of a Chemical Equation CH<sub>4</sub> (g) + 2O<sub>2</sub> (g) CO<sub>2</sub> (g) + 2 H<sub>2</sub>O (g) Stoichiometry Anatomy of a Chemical Equation Reactants appear on the left side of the equation. CH<sub>4</sub> (g) + 2 O<sub>2</sub> (g) CO<sub>2</sub> (g) + 2 H<sub>2</sub> ...

## Chapter 3 Stoichiometry—Chemistry

Stoichiometry is the tool for answering these questions. Stoichiometry The study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemi-cal reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass. Recall from Chapter 3 that the law states that

## Chapter 11: Stoichiometry

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## Chapter 03—Stoichiometry

Examples of Multiple Choice Questions from GENERAL CHEMISTRY. Choose your chapter: Fundamentals of Chemistry | Chemical Formulas & Composition Stoichiometry | Chemical Equations & Rxn Stoichiometry | Types of Chemical Reactions | | Atomic Structure | Chemical Periodicity | Chemical Bonding | Molecular Structure/Covalent Bonding Theories | Molecular Orbital Theory |

## Sample Exam Questions

Answer outline and marking scheme for question: 2. Molar mass of HgO = 201 + 16 = 217 gmol<sup>-1</sup>. 1.08g of HgO will contain 1.08 / 217 mols = 0.005mol From the equation, 1 mole of O<sub>2</sub>(g) is produced from 2 moles of HgO This means that 0.005 mol of HgO will produce 0.005 / 2 mol of O<sub>2</sub> = 0.0025 mol 0.0025 mol of O<sub>2</sub> will occupy 0.0025 x 24dm<sup>3</sup> = 0.06dm<sup>3</sup>

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