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Stoichiometry problems:

1. How many grams of silver chromate will precipitate when 150. mL of 0.500 M silver nitrate are added to 100. mL of 0.400 M potassium chromate? 2

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$\text{AgNO}_3(\text{aq}) + \text{K}_2\text{CrO}_4(\text{aq}) \rightarrow \text{Ag}_2\text{CrO}_4(\text{s}) + 2\text{KNO}_3(\text{aq})$

0.150 L  
 $\text{AgNO}_3$  0.500 moles  
 $\text{AgNO}_3$  1 moles  $\text{Ag}_2\text{CrO}_4$  331.74 g  
 $\text{Ag}_2\text{CrO}_4 = 12.4$  g  
 $\text{Ag}_2\text{CrO}_4$  1 L 2 moles ...

Solution Stoichiometry  
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Stoichiometry  
Worksheets with Answer  
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CHAPTER 9 REVIEW  
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SECTION 2

PROBLEMS Write the  
answer on the line to the  
left Stoichiometry  
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chapter 9. Show all your  
work in the space

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provided. 1. 4

Stoichiometry

worksheet answers

chapter 9. 5 mol The

following equation

represents a laboratory

preparation for oxygen

gas:  $2\text{KClO}_3(\text{s})$

$2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$  How

many moles of  $\text{O}_2$  form

if 3.0 mol of  $\text{KClO}_3$

are totally ...

Chapter 10 The Mole

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Solution

Name \_\_\_\_\_

CHEMISTRY 110 . last  
first ... Answer \_\_\_\_\_ 3]

60.5 mL of HNO<sub>3</sub> are  
required to react with

25.0 mL of a 1.00 M

Barium hydroxide

solution: 2 . HNO<sub>3</sub>(aq)

+ Ba(OH)<sub>2</sub>(aq) --> 2 .

H. 2. O(s) +

Ba(NO<sub>3</sub>)<sub>2</sub>(aq)

(BALANCED) Find the

. Molarity. of the nitric

acid solution . 25.0ml



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$\text{Ba}(\text{OH})_2 \times 1\text{L} = 10.3$   
 $\text{ml} \times 1\text{L} = 1.00\text{mol}$   
 $\text{Ba}(\text{OH})_2 \times 2\text{mol}$   
 $\text{HNO}_3 = 1\text{mol} \text{Ba}(\text{OH})_2$   
 $= 0.0500\text{mol HNO}_3$   
M ...

WORKSHEET 13  
Name - Cerritos College  
CHM 130  
Stoichiometry  
Worksheet The  
following flow chart  
may help you work  
*Page 33/38*

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stoichiometry problems.

Remember to pay careful attention to what you are given, and what you are trying to find. 1.

Fermentation is a complex chemical process of making wine by converting glucose into ethanol and carbon dioxide:  $C_6H_{12}O_6$  (s)  $\rightarrow$   $2 C_2H_5OH$  (l) +  $2 CO_2$  (g) A. Calculate the mass of ethanol

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CHM 130

Stoichiometry

Worksheet

Stoichiometry Practice  
Worksheet The correct  
answer is A.  $\text{H}_2\text{C}_2\text{O}_2$  is  
the limiting reagent:

$(4.5\text{mol H}_2\text{C}_2\text{O}_4)(2\text{mol MnSO}_4/5\text{mol}$

$\text{H}_2\text{C}_2\text{O}_4)$  According to  
the equation  $2\text{Na(s)} + 2$

$\text{H}_2\text{O(l)} \rightarrow 2$

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sodium reacts with  
water. Stoichiometry  
Flashcards | Quizlet

Worksheet 14 1

Worksheet #14 Limiting  
Reagents 1.

Stoichiometry Limiting  
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Answers ...

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