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| Solids**

**Properties, Crystalline**

**\u0026 Amorphous**

**, Lattice, Unit Cell. Solid**

**states | Defects in**

**solids (Schottky, Frenkel) |**

**Conductivity | Magnetic**

**Properties (All in 1)**

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**Intermediate II year |**

**Chemistry | Solid state |**

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**Chapter 1: Solid State (Part**

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**Solid State part 1**

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**Chemistry chapter 1 solid**

**state part -1 (L-1) Solid**

**states | Crystalline \u0026**

**Amorphous Solids | Types of**

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*Crystalline Solid | NEET JEE  
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Chapter 1: Crystal Structure  
Asst. Prof. Dr. Tagreed

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**Muslim 4 Basis:** An atom, ion, molecule or group of atoms is called a point and each point is connected with the other points to form a specific body. The basis for each point should be: 1 - Symmetrical parts in terms of structure, arrangement and direction. 2.

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If NaCl is doped with  $10^{-4}$  mol % of Sr Cl<sub>2</sub>, the concentration of cation vacancies will be ( $N_A = 6.02 \times 10^{23} \text{ mol}^{-1}$ ) [CBSE AIPMT 2007] A  $6.02 \times 10^{16} \text{ mol}^{-1}$

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electricity in solid state  
but ionic solids do so only  
in molten state or in  
solution or metals conduct  
electricity through  
electrons and ionic  
substances through ions.  
Metallic solids are  
malleable and ductile while  
ionic solids are hard and  
brittle.

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*Important Questions for  
Class 12 Chemistry Chapter 1  
The ...*

Hello, friends, this video contains a detailed explanation of Solid state intended for upcoming board exams. Please let me know if you have any doubts, Thank...

*Chemistry- Solid state class  
12 full chapter (Hindi ...*

Ans: It is a covalent or network solid. 1.8. Why are ionic solids conducting in the molten state and not in the solid state ? Ans: In the ionic solids, the electrical conductivity is due to the movement of the



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**Solutions:** Since the ionic mobility is negligible in the solid state, these are non-conducting in this state.

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