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Correct answer: I and III.

Explanation: Condition I is always true. Condition II is never true, as Gibbs free energy cannot be negative if

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enthalpy is positive and
entropy is negative.

Condition III may be true if
temperature is very high
(this is the scenario when
the term dominates the term.

Thermochemistry and

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Thermodynamics – Physical
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Problem : Given that the
free energy of formation of
liquid water is -237 kJ / mol ,
calculate the potential
for the formation of
hydrogen and oxygen from

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water. To solve this problem

we must first calculate ΔG

for the reaction, which is

$-2 (-237 \text{ kJ / mol}) = 474 \text{ kJ}$

$/ \text{ mol}$. Knowing that $\Delta G =$

$-nFE$ and $n = 4$, we

calculate the potential is

-1.23 V .

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disorder, is always highest

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