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: Crystal Field  
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Octahedral

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presents an up-to-date account of ligand field theory, the model currently used to describe the metal-ligand interactions in transition metal

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Compounds, and the way it is used to interpret the physical properties of the complexes.

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Theory describes the bonding, orbital arrangement, and other characteristics of coordination complexes. It represents an application of molecular orbital theory to transition metal complexes.



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A transition metal ion has nine valence atomic orbitals - consisting of five  $nd$ , one  $s$ , and three  $p$  orbitals. These orbitals are of appropriate energy to form bonding interaction with ligands. The LFT

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analysis is  
highly dependent  
on the geometry  
of the complex,  
but most  
explanations  
begin by de

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compounds, and  
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It examines the  
traditional  
electrostatic  
crystal field  
model, still  
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physicists, as well as covalent approaches such as the angular  
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the complexes.

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theory, in

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chemistry, one of several theories that describe the electronic structure of coordination or complex compounds, notably transition metal complexes, which consist of a central metal



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atom surrounded  
by a group of  
electron-rich  
atoms or  
molecules called  
ligands. The  
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theory explain  
some aspects of  
the chemistry of  
the transition  
metals, but  
neither model is  
good at  
predicting all  
of the  
properties of  
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complexes. A  
third model,  
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orbital theory,  
was therefore  
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Ligand field theory, in chemistry, one of several theories that describe the electronic structure of coordination or complex compounds,

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notably transition metal  
complexes, which  
consist of a  
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molecules called  
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developed in the  
1930s by two  
giants, Bethe  
and Van Vleck.

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Relating the  
Colors of  
Coordination

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Complexes And Its Spectrochemical Series According to the Crystal Field Theory, ligands that have high spin are considered "weak field" and ligands that have low spin are considered "strong field." This relates to

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the colors seen  
in a  
coordination  
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The ligand field  
theory is a firm  
background to  
foresee the

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magnetic properties of metallic complexes  $ML_n$  (M, transition metal ion; L, molecule or ligand). The ligand field splitting? oct between the energies of  $t_2g$  and  $e_g$  orbitals of an

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octahedral  
complex ML<sub>6</sub> is  
shown in Fig. 4

.

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The ligand field  
theory is a  
combination of  
both crystal



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theories. It was first proposed qualitatively by Griffith and Orgel. Ligand field theory is used to describe the bonding, orbital arrangement and other important

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Characteristics of coordination metal complexes.

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Crystal field theory treats interactions between the electrons on the metal and the

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Theory treats the metal-ligand interaction as a covalent bonding interaction, and depends upon considering the overlap between the d-orbitals on the metals and the ligand donor orbitals.

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In coordination chemistry, a ligand is an ion or molecule that binds to a central metal atom to form a coordination complex. The

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bonding with the metal generally involves formal donation of one or more of the ligand's electron pairs. The nature of metal-ligand bonding can range from covalent to ionic.

Furthermore, the

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metal-ligand bond order can range from one to three.

Ligands are viewed as Lewis bases, although rare cases are known to involve Lewis acidic "ligands".  
Metals and ...

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Model currently used to describe the metal-ligand interactions in transition metal compounds, and the way it is used to interpret the physical properties of the complexes.

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