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2009 HSC Chemistry Sample Answers . Question 20 (a) Sample answer:  $q = m \times C \times \Delta T$   $q = 210 \times 4.18 \times 65$  .  $q = 57057$  .  $q = 57.057 \text{ kJ}$  . Now  $\Delta H_{\text{comb}} = q/n$  .  $57.057$  .  $n \cdot C = 2$  .  $H$  .  $5$  .  $OH$  .  $1367$  .  $n = 0.04174 \text{ mol}$  .  $C$  .  $2$  .  $H$  .  $5$  .  $OH$  ?  $m \cdot C = n \times MM$  .  $2$  .  $H$  .  $5$  .  $OH$   $C$  .  $2$  .  $H$  .  $5$  .  $OH$  ?  $m \cdot C = 0.04174 \times 46.068$  .  $2$  .  $H$  .  $5$  .  $OH$  ?  $m \cdot C$  .  $2$  .  $H$  .  $5$  .  $OH = 1.923 \text{ g}$  . Now 50% of heat has been lost to surroundings.  $q_{\text{initial}}$  .  $m = 2 \times 1.920 \text{ g}$  .  $C$  .  $2$  .  $H$  .  $5$

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