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A3 Atomic structure 1. a) i) An atom is the smallest component of an element that can exist and still

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have the same chemical properties as the element. (1) ii) Mass number is the total number of protons and neutrons in the nucleus of one atom of an element.

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Question 6. Define the following terms: (a) Atomic number (b) Mass number (c) Nucleons (d) Valence shell Answer: (a) Atomic number: Atomic number refers to the number of protons present in an atom. It is denoted by Z. Example: An atom of oxygen contains 8 proton Therefore its atomic number is 8.

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3 Atomic structure 1. An atom is the smallest particle of an element that can exist by itself and still have the same chemical properties as the element. 2. - Protons - Neutrons - Electrons Protons and neutrons have a relative mass of 1, whereas electrons have a relative mass of 1/1840. Protons have a relative charge of +1, neutrons have no

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