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reversible reaction (q. 44). 19.2 In a  
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are continuously forming reactants. 2.1: A reversible reaction has reached a state of

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chromium com-pounds changes with the  
oxidation state of chromium FIGURE 19-2  
Sodium and chlo-rine react violently to form  
NaCl The synthesis of NaCl from its ele-ments  
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reversible reaction has reached a state of dynamic equilibrium. What does this information

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The  $\text{CaSiO}_3$  slag is less dense than the molten iron, so it can easily be separated. Also, the floating slag layer creates a barrier that prevents the molten iron from exposure to  $\text{O}_2$ , which would oxidize the Fe back to  $\text{Fe}_2\text{O}_3$ .

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amount of reactant changing per unit time ;  
3. Visible changes caused by chemical reactions are related to changes in the properties of

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22) For the reaction  $2 \text{C}_4\text{H}_{10} (\text{g}) + 13 \text{O}_2 (\text{g}) \rightarrow 8 \text{CO}_2 (\text{g}) + 10 \text{H}_2\text{O} (\text{g})$   $\Delta H^\circ$  is  $-125 \text{ kJ/mol}$  and  $\Delta S^\circ$  is  $+253 \text{ J/K} \cdot \text{mol}$ . This reaction is \_\_\_\_\_. A) spontaneous at all temperatures B) spontaneous only at high temperature C) spontaneous only at low temperature D) nonspontaneous at all temperatures E) unable to determine without more information

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