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Reading

Chapter 12 Stoichiometry127.

SECTION 12.1 THE ARITHMETIC

OF EQUATIONS (pages 353–358)

This section explains how to

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calculate the amount of reactants required or product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

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Answers:

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Stoichiometry 127. SECTION 12.1

THE ARITHMETIC OF EQUATIONS

(pages 353–358) This section

explains how to calculate the

amount of reactants required or

product formed in a nonchemical

process.

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Chapter 12 Guided Reading Stoichiometry Answer Key

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Stoichiometry127. SECTION 12.1
THE ARITHMETIC OF EQUATIONS
(pages 353–358) This section explains how to calculate the amount of reactants required or product formed in a nonchemical process.

Chapter 12 Stoichiometry Guided Reading Answers

Introduce the term stoichiometry in your own words.

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Stress that stoichiometry allows students to calculate the amounts of chemical substances involved in chemical reactions using information obtained from balanced chemical equations.

12.1 The Arithmetic of Equations

12

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Chapter 3: Stoichiometry –
Guided Reading Section 3.1 – 3.2
1. True or False? Most hydrogen atoms have a mass of 1.008 amu. Justify your answer. If true, explain why it is true. If false,

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What mass do most hydrogen atoms have? False, 1.008 amu is actually hydrogen's average mass, NO atom of hydrogen actually has the mass of 1.008 amu. 2.

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