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19. If 6.57 g of iron react with an excess of hydrochloric acid, HCl, then 11.2 g of iron(II) chloride are obtained in addition to hydrogen gas. Find

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the theoretical and percent yields.  $\text{Fe} + 2 \text{HCl} \rightarrow \text{FeCl}_2 + \text{H}_2$ . If 5.45 g of potassium chlorate are decomposed to form potassium chloride, 1.75 g of oxygen gas are also given off.

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(continued) Solve

each of the

following problems

as directed. Show

all your work. 13.

How many moles

Of magnesium are

required to react

with 2.0 mol of

hydrochloric acid?

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The equation for this reaction is  $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$ . 0 mo\ 14.

Aluminum reacts with HCl to produce aluminum chloride ( $\text{AlCl}_3$ ) and hydrogen ...

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statement is true,  
write "true." If it is  
false, change the  
underlined word or  
words to make it  
true. Write your  
answer on the line  
provided. 1. 2. 5.

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